

INCLUSION STUDIES ON α -CYCLODEXTRIN COMPLEXES OF GLIPIZIDE AND GLICLAZIDE WITH EFFECT OF PH

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ABSTRACT

Objective: To synthesize inclusion complexes of oral antidiabetic drugs, i.e., Glipizide (GP) and Gliclazide (Glc) with α -cyclodextrin (α -CD) and varying the pH which will enhance the solubility and thereby oral bioavailability of the drugs.

Methods: Liquid inclusion complexes were prepared and characterized by ultraviolet-visible (UV-VIS) spectroscopic studies, and fluorescent studies solid inclusion complexes of GP and Glc with α -CD were prepared by co-evaporation method and characterized by differential scanning calorimetric studies (DSC) and ^1H nuclear magnetic resonance spectroscopic studies (^1H NMR).

Results: The inclusion complexes of both GP and Glc with α -CD showed high binding constant values for pH 2 from the UV-VIS spectroscopy studies and fluorescent studies. DSC and ^1H NMR spectroscopic studies confirmed the inclusion of GP and Glc inside the α -CD cavity.

Conclusion: Formed with more stable inclusion complexes, with α -CD are GP and Glc at pH 2, along with pH modulation could be a suitable strategy to enhance the solubility and oral bioavailability of both the drugs.

Keywords: Solubility, Bioavailability, Glipizide, Gliclazide, Inclusion complexes, α -Cyclodextrin.

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INTRODUCTION

Poor solubility and oral bioavailability are the two great challenges in the pharmaceutical industry. Poor aqueous solubility hinders the *in vivo* efficacy of the drugs due to low bioavailability, abnormal pharmacokinetic profile, intersubject, interspecies variation [1] and thus poses challenging problems in the oral formulation development process.

Glipizide (GP), IUPAC name is 1-cyclohexyl-3-[[[2-(5-methylpyrazinecarboxamido) ethyl] phenyl]-sulfonyl] urea, Gliclazide (Glc) IUPAC name is 1-[4methylbenzene)sulfonyl]-3-octahydrocyclopenta[C] pyrrol-2-yl] urea.

GP and Glc are both second generation sulfonylurea hyperglycemic agents which are widely used in the treatment of noninsulin-dependent diabetes mellitus [2]. Both GP and Glc have very poor aqueous solubility and hence low bioavailability. Consequently, several studies have been developed for enhancement of solubility and bioavailability of GP and Glc through cyclodextrin (CD) inclusion complexes [3-7].

Enhancement in the solubility and the dissolution rate of such drugs can improve the oral bioavailability, which further improves the therapeutic efficacy and patient compliance. Different techniques have been used to improve the solubility and dissolution rate of the poorly soluble drugs, and in middle of the alternatives, the use of CD complexation presented a great interest [8] in recent years. CDs are cyclic oligosaccharides composed of glucopyranose units and can be represented as a truncated cone structure with a hydrophobic cavity. The hydrophobic cavity forms inclusion compounds with a wide range of guest molecules. Inclusion complex formation will modify the physicochemical properties such as solubility, stability, and bioavailability of poorly water-soluble drugs [9].

It is also a well-known fact that the pH of the medium is one of the important factors affecting the solubility of the drugs [10]. Hence, the present study is to investigate the effect of pH, in selected CD, i.e., α -CD

complexation on the solubility of GP and Glc. The solubility study is carried out in different media, i.e., pH 2 and 7 to identify the physiological pH at which the drugs have the highest solubility and also the influence of α -CD complexation in enhancing the solubility of GP and Glc.

METHODS

Reagents and materials

Analytical grade of GP, Glc, and α -CD from Sigma-Aldrich were used for the studies. The solvents used were of analytical grade. Triply distilled water was used for the preparation of stock solutions.

Instruments

- Systronics Smart Double beam Spectrophotometer-2203
- JASCO Spectrofluorometer FP-8200
- NETZSCH DSC 204 Calorimeter
- Bruker 300 MHz FT NMR Spectrometer.

Preparation of liquid inclusion complexes

About 0.0089 g of GP was accurately weighed and dissolved in 10 ml solution maintained at pH 2 and 7. Similarly, 0.0064 g of Glc was weighed and dissolved in 10 ml solution maintained at pH 2 and pH 7. About 0.2918 g of α -CD was weighed and dissolved in 30 ml distilled water. Inclusion complexes were prepared by varying the concentration of α -CD from 2×10^{-3} M to 1×10^{-2} M with both the drugs. The reactions were carried out at room temperature (303 K).

Preparation of solid inclusion complexes

Solid inclusion complexes of GP and Glc with α -CD were prepared by co-evaporation method.

Differential scanning calorimetry (DSC)

DSC measurements were performed on an NETZSCH DSC 204 calorimeter. Samples of 1-4 mg were accurately weighed in aluminum pans. These samples were then heated over a range of 25°C-300°C at a

constant rate of 10°C/minutes in a nitrogen purge of 50 ml/minutes. An empty aluminum pan was used as reference.

¹H nuclear magnetic resonance (¹H-NMR) spectroscopy

¹H-NMR spectroscopy studies were carried out using Bruker 300MHz FT NMR Spectrometer. Chemical shifts were reported in ppm (δ) downfield from tetramethylsilane (internal reference 0 ppm). Samples were prepared by dissolving in dimethyl sulphoxide.

RESULTS AND DISCUSSION

Effect of pH

Absorption studies

From Fig. 2a, it was observed that at pH 2, the absorption spectrum of GP was obtained at 230.8 nm in the absence of α -CD. After the addition of α -CD, there is a bathochromic shift from λ_{\max} 230.8 to 241.3 nm. The absorption spectrum showed a hypsochromic shift from λ_{\max} 235.5 to 229.5 nm as depicted in Fig. 2b. The spectral shifts were characteristic of the inclusion complexes formed by the encapsulation of the GP drug within the α -CD cavity.

For Glc, the absorption spectrum showed a bathochromic shift of λ_{\max} 225.9-235.7 nm indicating the inclusion complex formation at pH 2 as shown in Fig. 3a. A hypsochromic shift was observed with a λ_{\max} from 228.7 to 218.5 nm at pH 7 shown in Fig. 3b. The spectral shifts indicated the inclusion complex formation.

From Table 1, it was found that as the pH increases the intensity of absorption increases. The wavelength was shifted hypsochromically. After the addition of α -CD, a bathochromic shift was observed at pH 2, and a hypsochromic shift was observed at pH 7. This shows that as the pH of the medium increases, the solubility of both the drugs GP and Glc decreases.

Fluorescent studies

The fluorescent spectrum for GP α -CD complex exhibited a bathochromic shift of λ_{\max} 348-371 nm at pH 2 shown in Fig. 4a. As the pH was increased to 7, a hypsochromic shift was observed from λ_{\max} 387 to 369 nm as shown in Fig. 4b. The spectral shifts were indicative of encapsulation of the GP within the α -CD cavity leading to the formation of inclusion complexes.

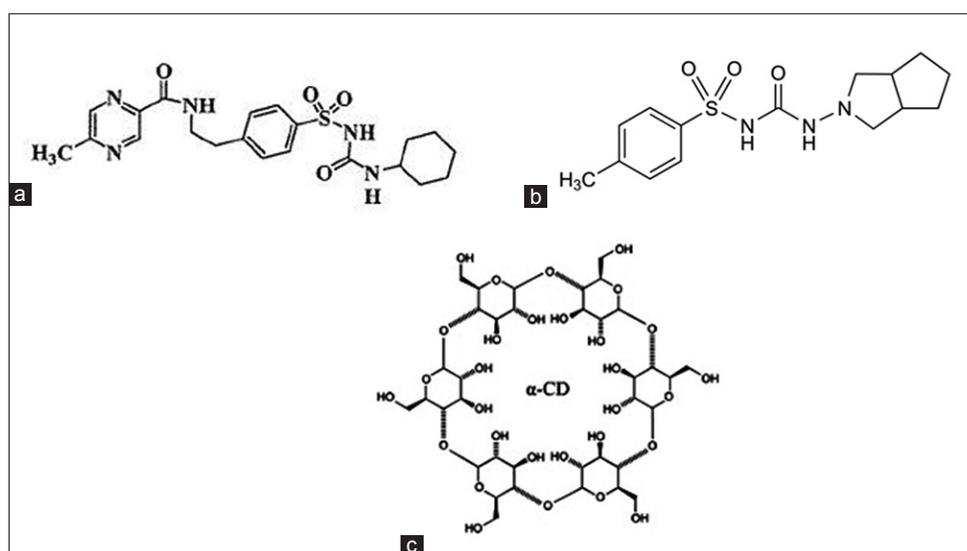


Fig. 1: Structure of (a) Glipizide (b) Gliclazide and (c) α -cyclodextrin

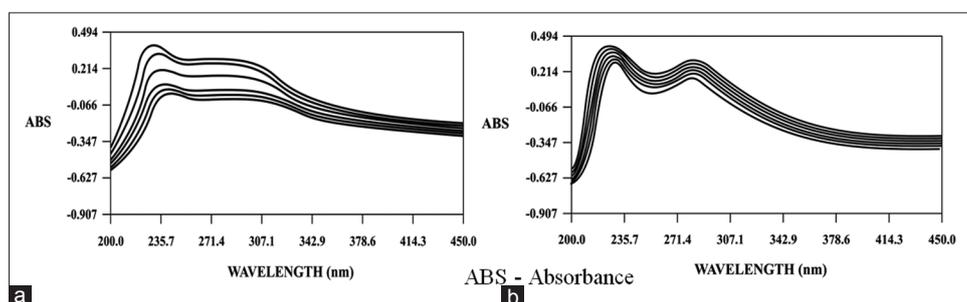


Fig. 2: Absorption spectrum of Glipizide- α -cyclodextrin at (a) pH 2 and (b) pH 7. ABS: Absorbance

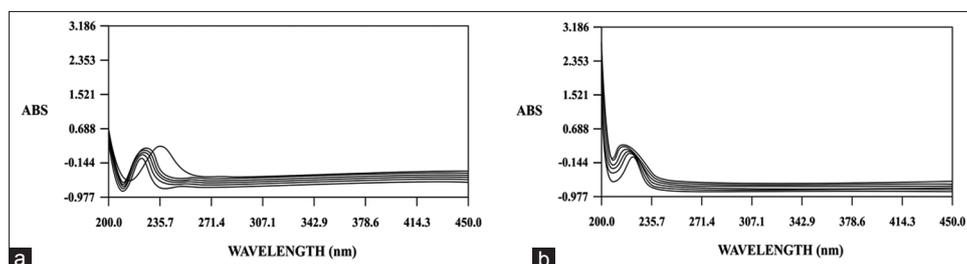


Fig. 3: Absorption spectrum of Gliclazide- α -cyclodextrin at (a) pH 2 and (b) pH 7

For Glc, a bathochromic shift was observed at pH 2 from λ_{max} 393 to 407 nm shown in Fig. 5a. At pH 7, a bathochromic shift was observed from λ_{max} 388 to 392 nm shown in Fig. 5b. The spectral shifts were characteristic of the inclusion complex formation between Glc and α -CD.

From Table 2, it was observed that, as the pH increases the intensity of emission decreases, the wavelength was shifted hypsochromically for GP α -CD complexes. For Glc α -CD complex, a bathochromic shift was observed with decrease in intensity. This depicted that as the pH of the medium increases the solubility of both GP and Glc decreases.

The α -CD dependence of GP on absorption can be analyzed by the Benesi-Hildebrand plot as given by [11]:

$$\frac{1}{A - A_0} = \frac{1}{A' - A_0} + \frac{1}{K_B(A' - A_0)[\alpha - CD]} \quad (1)$$

Where, $[\alpha$ -CD] represents the concentration of α -CD and A_0, A are the absorbances in the absence and presence of α -CD, respectively, and K_B is the binding constant.

From Fig 6a and 7a, it was observed that a reciprocal graph of $\frac{1}{A - A_0}$ versus $\frac{1}{[\alpha - CD]}$ yielded a straight line which indicated the formation of GP- α -CD inclusion complex with 1:1 and 1:2 stoichiometry. A reciprocal graph of $\frac{1}{A - A_0}$ versus $\frac{1}{[\alpha - CD]^2}$ yielded a curve which indicated the formation of Glc- α -CD inclusion complex with 1:2 stoichiometry.

Similarly, the α -CD dependence on the fluorescent spectra of GP could be analyzed by the Benesi-Hildebrand plot as given by:

$$\frac{1}{I - I_0} = \frac{1}{I' - I_0} + \frac{1}{K_B(I' - I_0)[\alpha - CD]} \quad (2)$$

Where, $[\alpha$ -CD] represents the concentration of α -CD and I_0, I are the intensities of fluorescence in the absence and presence of α -CD, respectively.

From Fig 6b and 7b, it could be inferred that a reciprocal graph of $\frac{1}{I - I_0}$ versus $\frac{1}{[\alpha - CD]}$ yielded a straight line which indicated the formation of GP- α -CD inclusion complex with 1:1 and 1:2 stoichiometry.

Table 1: The absorption spectral data of Glipizide and Gliclazide inclusion complexes at pH 2 and 7

Drugs	pH 2					pH 7				
	Concn. of α -CD	λ_{abs}	Abs	$\frac{1}{A - A_0}$	log ϵ	λ_{abs}	Abs	$\frac{1}{A - A_0}$	log ϵ	1/ $[\alpha$ -CD]
Glipizide	0	230.8	0.430		4.04	235.5	0.302		3.88	
	0.002	235.0	0.273	6.36	3.84	234.5	0.335	30.30	3.93	500
	0.004	236.4	0.209	4.52	3.72	233.5	0.343	24.39	3.94	250
	0.006	237.1	0.155	3.63	3.59	232.5	0.379	12.98	3.98	166.6
	0.008	239.3	0.117	3.19	3.47	231.5	0.385	12.04	3.99	125
	0.010	241.3	0.055	2.66	3.14	229.5	0.404	9.8	4.01	100
Gliclazide	0	225.9	0.103		3.43	228.7	0.096		3.40	
	0.002	227.3	0.132	34.48	3.54	226.6	0.132	27.7	3.54	500
	0.004	228.9	0.155	19.23	3.61	223.4	0.146	20.00	3.58	250
	0.006	229.1	0.164	16.39	3.63	221.0	0.171	13.33	3.65	166.6
	0.008	232.0	0.187	11.90	3.69	219.7	0.209	8.84	3.74	125
	0.10	235.7	0.262	6.289	3.83	218.5	0.226	7.69	3.77	100

Concn.: Concentration, Abs: Absorbance, Int: Intensity, A_0 : Absorbance of pure Glipizide/Gliclazide, A: Absorbance after addition of α -CD, $[\alpha$ -CD]: Concentration of α -CD, CD: Cyclodextrin

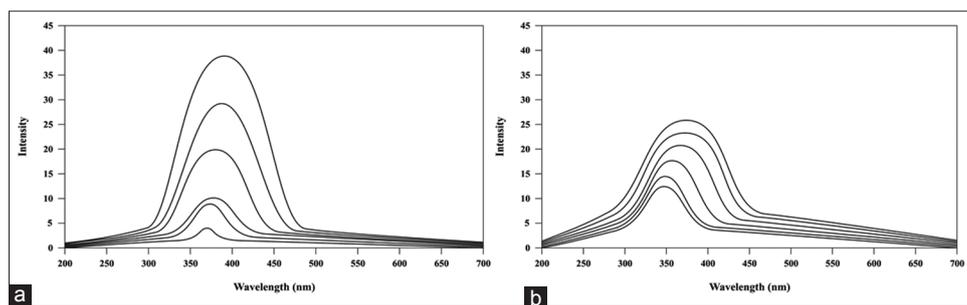


Fig. 4: Fluorescent spectrum of Glipizide- α -cyclodextrin at (a) pH 2 and (b) pH 7

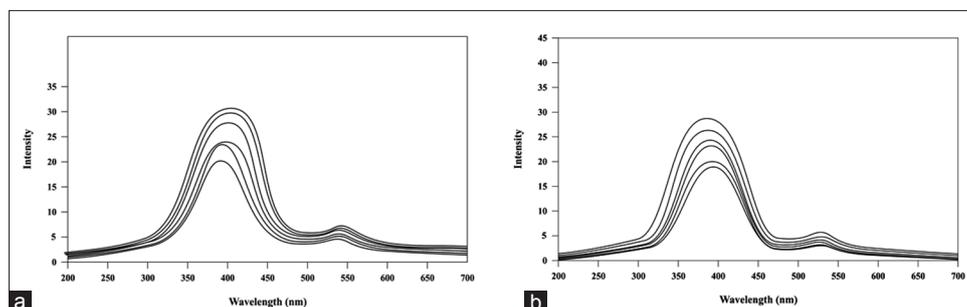
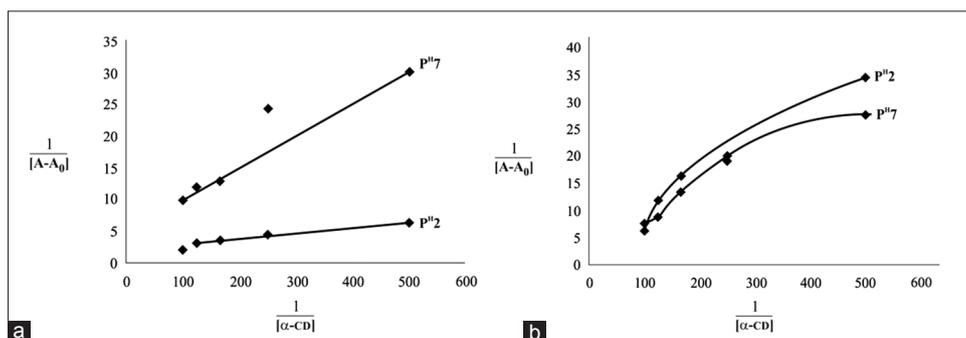
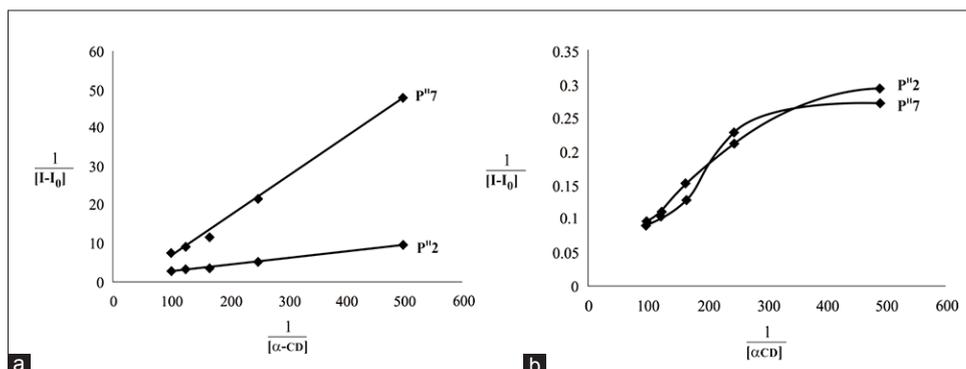


Fig. 5: Fluorescent spectrum of Gliclazide- α -cyclodextrin at (a) pH 2 and (b) pH 7

Table 2: The fluorescent spectral data of Glipizide and Gliclazide inclusion complexes at pH 2 and pH 7

Drugs	pH 2			pH 7				
	Concn. of α -CD	λ_{flu}	Int.	$\frac{1}{I-I_0}$	λ_{flu}	Int.	$\frac{1}{I-I_0}$	$1/[\alpha\text{-CD}]$
Glipizide	0	348	12.46		387	38.64		
	0.002	349	14.54	0.480	385	28.32	0.0968	500
	0.004	354	17.08	0.216	380	19.14	0.0534	250
	0.006	361	21.11	0.115	378	10.15	0.0351	166.6
	0.008	368	23.4	0.091	373	8.86	0.0333	125
	0.010	371	25.7	0.071	369	4.2	0.0290	100
Gliclazide	0	393	20.03		388	29.4		
	0.002	395	23.7	0.2724	389	26.0	0.2941	500
	0.004	398	24.4	0.2288	390	24.7	0.2127	250
	0.006	401	27.8	0.1287	391	22.9	0.1538	166.6
	0.008	404	29.7	0.1034	392	20.1	0.1075	125
	0.010	407	31.0	0.0911	392	18.9	0.0952	100

λ_{flu} : Wavelength of fluorescence, Int.: Intensity, I_0 : Intensity of drug without α -CD, I: Intensity after addition of α -CD, $[\alpha\text{-CD}]$: Concentration of α -CD, CD: Cyclodextrin

Fig. 6: Benesi Hildebrand plot of (a) Glipizide- α -cyclodextrin (CD) at pH 2 and 7 (b) Gliclazide- α -CD at pH 2 and pH 7Fig. 7: Benesi Hildebrand plot of (a) Glipizide- α -cyclodextrin (CD) at pH 2 and 7 (b) Gliclazide- α -CD at pH 2 and pH 7

A reciprocal graph of $\frac{1}{I-I_0}$ versus $\frac{1}{[\alpha\text{-CD}]^2}$ yielded a straight line which indicated the formation of Glc α -CD inclusion complex with 1:2 stoichiometry.

A linear relation is obtained in the case of absorption and fluorescence at both pH values for GP and Glc which revealed the stoichiometry of the complex formed by the GP with α -CD is 1:1,1:2 and that of Glc with α -CD is 1:2. Moreover, the linearity relation also indicated the increase in aqueous solubility of the drug upon complexation with increase in the concentration of α -CD.

The binding constant for GP α -CD complex from absorption studies was found to be 266/M at pH 2 and 245/M at pH 7. From fluorescent studies, it was evaluated as 276/M at pH 2 and 80.97/M at pH 7.

For Glc α -CD complexes, the binding constant was calculated to be 117.5/M at pH 2 and 85.47 at pH 7 from absorption studies. Fluorescent

studies showed that the binding constant was 112.0/M at pH 2 and 84.1/M at pH 7.

As the value of binding constant increases, the stability of the inclusion complex also increases and hence the solubility of the drug increases. It was found that the value of binding constants decreases with increase in pH from 2 to 7. So, with increase in pH, the stability of the inclusion complexes formed between GP and Glc with α -CD and hence the solubility of both GP and Glc decreases.

DSC studies

DSC analysis was performed to characterize inclusion complexes with α -CD by comparing the thermal behavior of the individual components as well as their individual components [12-14].

When the guest molecules are incorporated into the CD cavity or in the crystal lattice, their melting point and boiling point usually shift to different temperatures or disappear within the temperature range

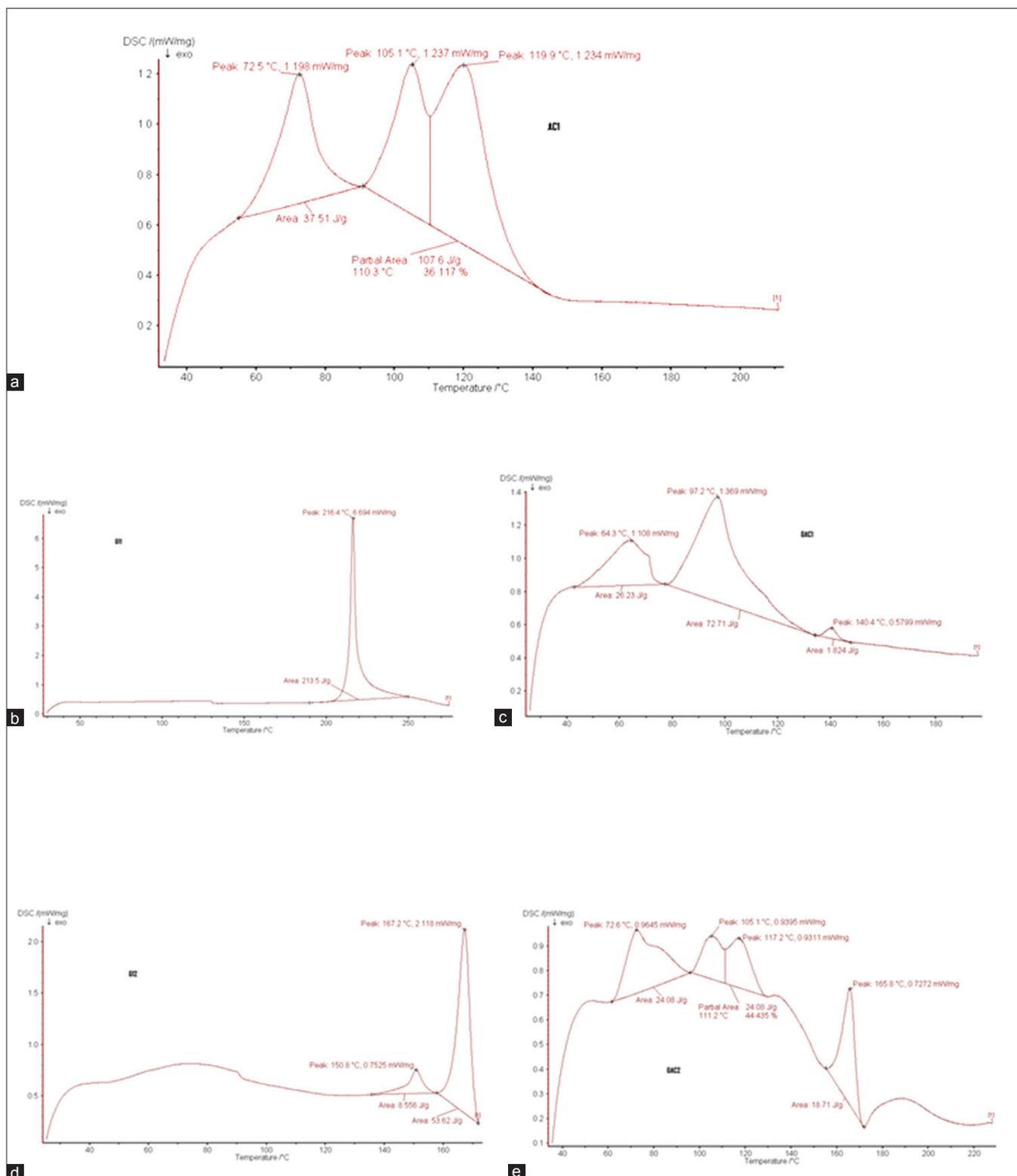


Fig 8: Differential scanning calorimetric thermograms of (a) α-cyclodextrin (CD) (b) Glipizide (c) Glipizide α-CD (d) Gliclazide and (e) Gliclazide α-CD complex

when CD lattice is decomposed. This is one of the best method for detecting drug-CD complexation [15]

The DSC thermograms of α-CD showed three exothermic peaks at 72.5°C, 106.1°C, and 119.9°C which correspond to the dehydration of α-CD (Fig. 8a). Further, the thermogram was broadened indicative of the decomposition of α-CD and transition from crystallinity to amorphous

nature. The DSC thermogram of GP showed one sharp characteristic exothermic peak at 216.4°C and 10.4°C corresponding to its melting point (Fig. 8b). However, in the case of the GP α-CD thermograms, three exothermic peaks were observed at 64.3°C, 97.2°C, and 140.4°C corresponding to the polymorphs formed during the fusion process via the dehydration of α-CD. The peak corresponding to the melting point of GP became extinct or disappeared in the DSC thermogram (Fig. 8c).

The extinction or disappearance of the exothermic peak of GP in GP α -CD thermogram was strong evidence for the total encapsulation of the GP in the α -CD cavity. This also suggested the formation of 1:1 stoichiometry of inclusion complex between α -CD and GP.

The DSC thermograms of Glc showed a sharp characteristic exothermic peak at 167.2°C corresponding to its melting point (Fig. 8d). The broad peak around 40°C-80°C is due to the amorphous nature of Glc. The DSC thermogram of Glc α -CD complex indicated the characteristic exothermic peaks found in α -CD, i.e., at 72.6°C, 105.1°C, and 117.2°C due to the dehydration of α -CD (Fig. 8e).

Moreover, there is a shift of the exothermic peak characteristic of Glc from 167.2°C to 165.8°C, toward lower temperature. This is indicative of the partial encapsulation of the Glc drug in the α -cyclodextrin cavity. Hence, the inclusion complex formed is not of type 1:1. Instead, the stoichiometry was found to be 1:2.

NMR spectroscopic studies

The most direct evidence for the inclusion of a guest molecule into a CD cavity is obtained by $^1\text{H-NMR}$ spectroscopy [16,17]. $^1\text{H-NMR}$ may also be used to determine the direction of penetration of guest molecule into the CD cavity [18,19]. The H-3 and H-5 protons of CD which are directed toward the interior of the CD will show a significant upfield shift if inclusion does indeed occur and the H-1, H-2, and H-4 protons located at the exterior of the cavity will show marginal upfield shift. The spectrum of the guest molecule may also be changed upon inclusion complex formation [20,21].

The $^1\text{H-NMR}$ chemical shift for each proton of α -CD was evaluated for the formation of inclusion complex between GP and α -CD.

The formation of inclusion complexes was studied by the changes in the chemical shifts caused by the guest and the host to each other [22-24]. In the structure of α -CD, H-3 and H-5 protons are located inside the cavity, H-3, H-5 located near the wider rim, and H-5 located at the narrow rim of the CD cavity. All other protons H-1, H-2, and H-4 are located on the exterior of the cavity.

From the Table 3, $\Delta\delta\text{H-3}=\Delta\delta\text{H-5}=-0.17<0$, for the GP- α -CD complex. This showed that inclusion or encapsulation of GP inside the α -CD cavity took place, and an upfield was observed in the inclusion complex due to the dissolution of GP inside the α -CD cavity.

Further, the spectrum of pure GP was varied in the case of GP α -CD complex. This proved as a strong evidence for the inclusion complex formation.

The observed shielding effect can be attributed to the diamagnetic anisotropy effect due to the inclusion of a group rich in π electrons (aromatic ring) into the hydrophobic central cavity of α -CD molecule.

From Table 4, it was observed $\Delta\delta\text{H-3}>\Delta\delta\text{H-5}>0$, which indicated encapsulation of the drug Glc in the α -CD cavity took place. The upfield shift observed in Glc α -CD inclusion complex was due to the dissolution of Glc drug inside the α -CD cavity and are shielded by the H-3 and H-5 protons of α -CD.

Comparative study of the effect of pH on Glipizide and Gliclazide α -CD inclusion complexes

From the absorption and fluorescent studies, it was observed that the binding constant values were higher at pH 2 for both GP and Glc α -CD inclusion complexes. Ionized species were present at pH 2 in both the complexes. However, it may be assumed that the ionic species were sufficiently far away from the hydrophobic cavity of the α -CD. The binding constant values of GP and Glc α -CD complexes were comparatively lower at pH 7. This phenomenon was exhibited by both the drugs because of the sulphonylurea group present in both the drug. The other groups do not interfere much with inclusion

Table 3: ^1H -chemical shifts (ppm) of the protons of α -cyclodextrin and Glipizide in the free and inclusion complex states

Proton α -CD	δ_0 (free) α -CD	δ_c (complex) Glipizide- α -CD	$\Delta\delta$ ($\delta_c-\delta_0$)
H-1	4.98	4.97	-0.01
H-2	3.63	3.62	-0.01
H-3	3.82	3.99	-0.17
H-4	3.40	3.77	-0.03
H-5	3.75	3.58	-0.17
H-6	3.92	3.89	-0.03

δ_0 : Chemical shift of pure α -CD, δ_c : Chemical shift of Glipizide α -CD complex, $\Delta\delta$: Difference in chemical shift between complex and pure α -CD. CD: Cyclodextrin

Table 4: $^1\text{H-NMR}$ chemical shifts (ppm) of the protons of α -CD and Gliclazide in the free and inclusion complex states

Protons α -CD	δ_0 (free) α -CD	δ_c Gliclazide α -CD	$\Delta\delta=\delta_c-\delta_0$
H-1	4.98	-	-
H-2	3.63	3.27	-0.36
H-3	3.96	2.78	-1.18
H-4	3.40	2.62	-0.78
H-5	3.78	2.47	-1.31
H-6	3.92	2.33	-1.59

δ_0 : Chemical shift of pure α -CD, δ_c : Chemical shift of Gliclazide α -CD complex, $\Delta\delta$: Difference in chemical shift between complex and pure α -CD. CD: Cyclodextrin

complex formation. Increase in pH of the medium decreases the stability of the inclusion complexes formed and hence the solubility of both GP and Glc. Therefore, it was evident that the most favorable pH for increasing the solubility of the drugs was at pH 2 and less favorable pH was 7. It was often assured that the unionized species were primarily responsible for the formation of the complex and subsequent increase in solubility. However, ionization of a drug can also result in enhanced complexation, and the solubility of the ionized drug complex can sometimes exceed the solubility of the unionized drug complex α -CD plays an important role in the solubility of GP and Glc together with pH adjustment. The solubility of the corresponding complexes can be much higher than what could be expected by increasing the total drug solubility through only appropriate pH adjustment. This is possible due to the specific interaction of the drugs molecules with the hydrogen bond system of the α -CD and/or the modification of the hydrogen bonding network of the surrounding water molecules, the drugs can be strongly bound inside the α -CD cavity, even when they are in ionized form because their basic center, and hence, the charge can be located quite far away from the molecules hydrophobic part that is included in the α -CD cavity so that the complexation ability is not significantly compromised.

CONCLUSION

The absorption and fluorescent spectral studies confirmed the formation of α -CD inclusion complexes with both GP and Glc. From absorption and fluorescent studies, stoichiometry of GP α -CD complex is 1:1, 1:2 and that of Glc α -CD complex is 1:2. ^1H NMR and DSC studies also served as evidence for the formation of inclusion complexes. The soluble and stable complexes were formed at pH value 2. The present study has demonstrated the possibility to improve the aqueous solubility of GP and Glc by simultaneous α -CD complexation and pH value modulation. High stability constant values were obtained only at pH 2 rather than pH 7. This has suggested that α -CD complexation accompanied with pH value modulation may be a suitable strategy to enhance the aqueous solubility of both GP and Glc, thus overcoming the problems associated with low bioavailability in the pharmaceutical formulation of both the drugs.

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