GREEN SYNTHESIS OF SUPERPARAMAGNETIC IRON OXIDE NANOPIRATE FROM FICUS CARICA FRUIT EXTRACT, CHARACTERIZATION STUDIES AND ITS APPLICATION ON DYE DEGRADATION STUDIES

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ABSTRACT

Objective: The synthesis of nanoparticles (NPs) has become a matter of great interest in recent times due to their various advantageous properties and applications in a variety of fields. Metal NPs are being increasingly used in many sectors, and there is growing interest in the biological and environmental safety of their production.

Methods: In this study, iron oxide NPs (Fe₃O₄-NPs) were synthesized from fruits of Ficus carica using a rapid, single step and completely green biosynthetic method by reduction of ferrous sulfate solution with F. carica ethanolic extract. The prepared Fe₃O₄-NPs were investigated by X-ray diffraction, Fourier transform infrared spectroscopy, and ultraviolet-visible spectroscopy.

Results: The report emphasizes the effect of superparamagnetic Fe₃O₄-NPs on the degradation rate of hazardous dyes acid blue.

Conclusion: To conclude, Fe₃O₄-NPs were prepared from fruits of F. carica using a rapid, single step and completely green biosynthetic method by reduction of ferrous sulfate solution with F. carica ethanolic extract.

Keywords: Ficus carica, Ethanolic extract, Reduction, Ferrous sulfate, Superparamagnetic iron oxide nano particles, Dye degradation.

INTRODUCTION

Nanotechnology is a multidisciplinary branch of science that encompasses numerous fields of science and technology, including biomedicine, pharmaceutics, environmental science, and others [1]. The use of nanoparticle (NP) materials offers advantages due to their unique size and physicochemical properties. NPs are clusters of atoms in the size range of 1-100 nm. The control of the monodisperse size of NP is very important because the properties of the nano crystals strongly depend on the dimension of the NPs. One of the most commonly-used nanoscale materials is magnetic NPs (MNP)s: A type of core/shell NP structure that consists of a magnetic core encapsulated in an organic or a polymeric coating. Without a coating, MNPs have hydrophobic surfaces with large surface-to-volume ratios and a propensity to agglomerate [2]. Iron oxide NPs (Fe₃O₄-NPs) have attracted intensive research interest because of their important applications in cancer therapy, drug delivery, magnetic resonance imaging (MRI), and wastewater treatment [3]. Superparamagnetic Fe₃O₄-NPs with appropriate surface chemistry can be used for numerous in vivo applications, such as MRI contrast enhancement, tissue repair, and immunomodulation, detoxification of biological fluids, hyperthermia, drug delivery, and cell separation [4]. The biosynthesis of Fe₃O₄-NPs of different sizes and shapes has been reported using bacteria [5,6] and plant extract [7]. The stability of iron NPs against aggregation can be improved by imparting electrostatic repulsion, applying organic surfactants, or through the use of capping agents [8]. In recent times, sources of dye contamination have expanded from textile industries to food, paper, printing, cosmetic, and pharmaceutical companies [9]. When these dyes are not treated properly, they get accumulated in the environment and become a threat to the ecosystem. These non-treated dyes are potentially carcinogenic, mutagenic and genotoxic, example of such dyes include Acid Red 26, Direct Blue 6, Direct Black 38, etc. Abatement of dyes is a required part of wastewater treatment. Nanotechnology has been extended to the wastewater treatments in the recent years. Due to high surface area silver NPs exhibits an enhanced reactivity [10,11].

Ficus carica is known to contain polyphenols and flavonoids that act both as a reducing agent and a capping agent. Its fruit, root, and leaves are used in traditional medicine to treat various ailments such as gastrointestinal (colic, indigestion, loss of appetite, and diarrhea), respiratory (sore throats, coughs, and bronchial problems), and cardiovascular disorders and as anti-inflammatory and antispasmodic remedy.

A wide variety of methods have been reported in for the synthesis of Fe₃O₄-NPs such as microemulsion technique [12], electrochemical route [13], hydrothermal process [14], sonochemical method [15], and co-precipitation method [16]. Nowadays, synthesis of Fe₃O₄-NPs using phytochemicals has attracted much attention due to their simplicity, environmental benignity, and low cost.

Dyes are a major class of synthetic organic compounds released by many industries such as paper, plastic, leather, food, cosmetic, textile, and pharmaceutical industries [17,18]. The synthesis and applicability of the Fe₃O₄-NPs in the effective removal of acid blue have been reported. The main objective of this study is to test the applicability of Fe₃O₄-NPs as a catalyst in the removal of acid blue dye. Photocatalytic degradation of acid blue was conducted using Fe₃O₄ NP as a catalyst in the presence of sunlight.

METHODS

Sample collection
Fig fruits were collected from Anna University, ACT Campus, Chennai, Tamil Nadu, India.
Preparation of ethanolic extract
Fig fruits were shade dried and grounded well. Ethanol (70%) was used as a hydro-alcoholic solution for soaking. 15 g of grounded F. carica fruit was soaked in 200 ml of 70% ethanol in a glass beaker for a day. Beaker was sealed with aluminum foil and kept in a magnetic stirrer. For coarse filtration, soaked material was filtered several times using filter paper. This coarse filtrate was then filtered through a Whatman filter paper. Using hot air oven, the above filtrate was evaporated at 50°C under reduced pressure until the concentrate was reduced to 1/3rd of the initial volume. Dark brown colored extract with residues was obtained. The extract, which obtained was then again filtered to remove residues, final extract obtained was dark brown in color and it was stored in a refrigerator.

Preparation of Fe₃O₄-NPs
In a typical reaction procedure, 20 ml of F. carica ethanolic extract was added to 270 mg of FeSO₄ and pH is adjusted to 9, using 0.1 M NaOH solution, then the solution was placed under vigorous magnetic stirring for 4 hrs at 80°C. During this process, the color of the reaction solution changed from yellowish translucent to a brown color, indicating the formation of Fe₃O₄-NPs. The resulting product, iron NP was centrifuged at 7000 rpm for 15-20 minutes and washed several times with 1:1 mixture of distilled water, and absolute methanol. The purified NPs powder was dried at 90°C for 16 hrs and stored in an airtight bottle for further characterization by UV, scanning electron microscope, Fourier transform infrared spectroscopy (FTIR), X-ray diffractometer (XRD), and vibrating sample magnetometer.

Procedure for dye degradation
To 10 ml of dye solution, Fe₃O₄-NPs were added and the suspension was subjected to irradiation. Experiments were carried out under sunlight. The aqueous suspension was magnetically stirred throughout the experiment. At different time intervals, aliquot was taken out, and the absorption spectra were recorded and rate of decolorization was observed in terms of change in intensity at λ max of the dyes. The decolorization efficiency % has been calculated as:

\[
\text{Efficiency} \% = \frac{C_{0} - C}{C_{0}} \times 100
\]

Where C₀ = Initial concentration of dye and C = Concentration of dye after photo irradiation.

RESULTS AND DISCUSSION

UV-VIS spectroscopy
The strong interaction of metal NPs with light results in the collective oscillation of the conduction electrons on the metal surface, known as a surface plasmon resonance (SPR). The SPR results in unusually strong scattering and absorption properties. Due to the unique optical properties of Fe₃O₄-NPs, a great deal of information about the physical state can be obtained by analyzing the spectra. As the diameter increases, the peak plasmon resonance shifts to longer wavelengths and broadens. Ultraviolet-visible (UV-VIS) spectrum of Fe₃O₄-NPs is shown in the Fig. 1. Here, the sharp peak at 283 nm indicates the presence of Fe₃O₄-NPs. Parallel results were observed by Awwad and Salem [19]. Phytochemicals in the F. carica fruit extract was shown to reduce the Fe₃O₄-NPs present in the fruit extract. The peak at 2367/cm is due to the stretching vibration of C-H derived from aromatic rings that are present in the extract. The peak at 2367/cm and 2344/cm is due to the stretching vibration in aromatic. The strong peak at 1245/cm is due to the stretching vibration of C=O bond in the alkenes ring. The peak at 1158/cm is due to C-O stretching in aldehydes. The peak at 1158/cm is assigned to C=C ring stretching vibration in Aromatic. The strong peak at 1245/cm is due to the stretching vibration of C=O bond in the alkenes ring. The peak at 1158/cm is due to C-O stretch. The peak at 1158/cm is assigned to C-F stretch in carboxylic acid, ester. The peak at 1158/cm is due to C-H out of plane (Bending moles). From the FTIR data, it is clear that the bioactive molecules present in the leaves extract of F. carica interacted with the synthesized Fe₃O₄ NPs.

XRD
The typical powder XRD pattern of the prepared nanoparticles is shown in Fig. 3. XRD analysis can give you the idea of crystallinity, purity, and size of your synthesized metal NPs. Data were taken for the 2θ range of 10°-70°. The peaks at 31.84°, 35.88°, 45.56°, and 56.88° are assigned to diffraction from the (1 0 0), (3 1 1), (3 3 1), and (3 3 3) planes of face-centered cubic crystal. The observed peaks were compared with the standard powder diffraction card of (JCPDS File No. 87-0720), and the results are in agreement with the standard XRD pattern of Fe₃O₄ NPs.

Particle size calculation
From this study, considering the peak at degree, average particle size has been estimated using Debye-Scherrer formula [13,15].

\[
D = \frac{0.94 \lambda}{β \cosθ}
\]

Where D = Particle size, λ = X-ray wavelength, β = Half width at half maximum, θ = Bragg’s angle.
Where "\( \lambda \)" is wave length of X-ray (0.1541 nm), "\( \beta \)" is full width at half maximum, "\( \theta \)" is the diffraction angle, and "D" is particle diameter size.

i. \( 2\theta = 31.84 \)

Catalysis occurs only on the surface of metals, and hence increasing the available surface area of the NP will greatly enhance the effectiveness of the catalyst [21]. Decreasing the particle size will increase the catalytic activity, but there is a critical size below which proves that further decreases will hinder the catalytic activity [22]. Metal NPs help in the electron transmits from donor to the acceptor. NPs possess a large surface area which acts as a substrate for the electron transfer reaction. Just before the electron transfer reaction, both of the reactants are adsorbed on the metal surface. Subsequently, the reactant gains an electron and is reduced (Figs. 4 and 5).

### CONCLUSION

To conclude, Fe\textsubscript{3}O\textsubscript{4}-NPs were prepared from fruits of F. carica using a rapid, single step, and completely green biosynthetic method by reduction of ferrous sulfate solution with F. carica ethanolic extract. The prepared Fe\textsubscript{3}O\textsubscript{4}-NPs were characterized by X-ray diffraction, Fourier transform infrared spectroscopy, and UV-VIS spectrophotometer. The preliminary work confirms the effect of superparamagnetic Fe\textsubscript{3}O\textsubscript{4}-NPs on the degradation rate of hazardous dyes acid blue.

### REFERENCES


